



CHEMISTRY

CH: 2 Acids, Bases and Salts

Name: _____

Date: _____

Class: XSec: ____

1. What is chlor-alkali process? What are the products obtained in this process? Give their uses.
2. Compound A when dissolved in water gives compound B which is used in whitewashing. Compound B reacts with CO_2 to form a white precipitate of compound C. Identify compounds A, B & C. Also write the chemical equations involved
3. Acidic character is shown only in aqueous solution. Describe an activity to prove it.
4. What would be the colour of litmus in a solution of sodium carbonate?
5. How would you distinguish baking soda and washing soda by heating?
6. A solution reacts with crushed egg shells to give a gas that turns lime water milky. What is the nature of solution? Which is the gas evolved?
7.
 - a) What do you understand by pH?
 - b) There are two solutions A and B. Solution A turns pink in a few drops of phenolphthalein is added to it. Solution B changes the colour of blue litmus to red. What range of pH would these solutions have?
8. Dry pellets of a base 'X' when kept in open absorbs moisture and turns sticky. The compound is also formed by chlor-alkali process. Write chemical name and formula of X. Describe chlor-alkali process with balanced chemical equation. Name the type of reaction occurs when X is treated with dilute hydrochloric acid. Write the chemical equation.
9. Name the acid present in the following: lemon, tomato, orange, vinegar, tamarind, curd.
10. How would you distinguish between baking powder and baking soda by heating?
11. Give suitable reason for the following
 - a) The crystals of washing soda change to white powder on exposure to air.
 - b) An aqueous solution NaCl is neutral but an aqueous solution of Na_2CO_3 is basic.
12. Name the salt used to remove the permanent hardness of water.



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Copper sulphate crystals lose their blue colour and turn white while heated strongly. Give reason.

14. At what pH in the mouth is tooth decay faster and why?

15. Arrange the following in the increasing order of pH.

Milk of magnesia, lemon juice, pure water, sodium hydroxide.

16. A yellow powder X gives a pungent smell if left open in air. It is prepared by the reaction of dry compound Y with chlorine gas. It is used for disinfecting drinking water. Identify X and Y and write the reaction involved

17. You might have seen lemon or tamarind juice being used to clean tarnished surface of copper vessels. Explain why these sour substances are effective in cleaning vessels.

18. If acetic acid & hydrochloric acid of same concentration are taken, which of the two is a stronger acid and why?

19. Give reasons for the following:

- a) Aluminium is a reactive metal but it is still used for packing food particles.
- b) Calcium starts floating when water is added to it.
- c) Eating sweets leads to tooth decay.

20. What are strong and weak acids? In the following list of acids, separate strong and weak acids; HCl, H_2CO_3 , H_2SO_4 , HNO_3 , Oxalic acid, formic acid.